**Gas Laws Quiz– Honors**

1. Gases behave least ideally at low temperatures and most ideally at high temperatures. Given what you know of the kinetic molecular theory, explain why this might be. (5 pt)
2. Why is it reasonable to say that the molecules of ideal gases are infinitely small? (5 pt)
3. If I move a 2.5 L balloon from the surface of a pool (pressure = 1.00 atm) to 1 meter under the water (pressure = 1.03 atm), what will its new volume be? (5 pt)
4. A burp usually has a volume of about 0.015 L. Assuming the body temperature of a human is 37 degrees Celsius and the pressure inside the stomach is about 1.025 atm, how many moles of gas are in a typical burp? R = 0.08206 Latm/molK (5 pt)
5. What is an ideal gas? Give an example, if appropriate. (3 pt)